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1	Solar photocatalysis of TiO ₂ supported natural palygorskite nanofibers elaborated by a
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23 Abstract:

This study reports the successful synthesis of supported TiO₂_Palygorskite nanocomposites by 24 a one-pot dry mechanochemical route. Indeed, the elaboration procedure involved an in-situ 25 26 reaction between accessories carbonates present in raw fibrous palygorskite clay and titanyl sulfate (TiOSO₄) precursor under variable grinding conditions, essentially ball/solid matter 27 mass ratio and rotation velocity. This yielded after air annealing at 600°C for 1 h to the 28 immobilization of anatase TiO₂ nanoparticles (≈ 8 nm of average size) as evidenced by XRD 29 and TEM analyses. Once the conditions of elaboration were optimised, the photocatalytic 30 properties were evaluated under 3 conditions: artificial UV radiation, artificial solar radiation 31 (UV+visible range) and under dynamic solar illumination taking into account the 32 33 discontinuities of the solar resource. The results allowed the estimation and comparison of the catalyst's capabilities and showed its ability to work under natural irradiation. The so developed 34 supported photocatalysts TiO₂/Palygorskite exhibited a good activity towards the removal of 35 36 Orange G (OG) dye from aqueous media under artificial UV and natural solar radiations.

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Keywords: TiO₂, palygorskite, mechanochemical synthesis, nanocomposite, solar
photocatalysis, wastewater treatment.

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46 **1. Introduction**

Heterogeneous photocatalytic oxidation recently has emerged as an efficient alternative process 47 48 for wastewater treatment [1-6]. The principle of this technique relies on the creation of reactive species as holes (h⁺) and hydroxyl radicals (OH[•]) upon irradiating a semiconductor oxide with 49 an energy source (h υ) higher than its energy band gap [1,6]. In optimized processes, the highly 50 51 oxidizing species so generated are able to induce complete mineralization of organic pollutants into CO₂ and H₂O [1]. TiO₂ anatase is the most active semiconductor oxide in photocatalysis 52 and is besides widely used owing to its numerous advantages, for instance, non-harmfulness, 53 low cost, and chemical inertness [1]. Several wet methods, including sol-gel [7–9], flame 54 aerosol [10], solvothermal [11] were reported for the synthesis of TiO₂ powder. Interestingly, 55 56 a dry method based on grinding powders of carbonate compound, namely Na₂CO₃ and a Ti molecular precursor such as titanium sulfide (TiS₂) [12,13] and titanyl sulfate ((TiOSO₄ $2H_2O$) 57 [14-16] was also reported to synthesize pure TiO₂ powder. However, the direct use of TiO₂ in 58 59 the form of nano-powder (e.g., commercial Degussa P25 powder [10]) raises several problems such as agglomeration of the nanoparticles during the process, which reduces photocatalytic 60 efficiency [17]. Additionally, recovering of micron sized aggregated particles from water 61 decontaminated by TiO₂ slurry needs to implement costly microfiltration processes [2,6,17,18]. 62 To overcome these drawbacks, researches focused on improving photocatalytic processes by 63 the development of TiO₂ supported photocatalysts, in particular starting with natural highly 64 dispersed materials as support. 65

The solar resource is a sustainable energy by definition and available for many applications. For photo-oxidation applications, capacities are limited because most catalysts are only sensitive to ultra violet (UV) radiation in the spectrum. This represents only 5% of the solar radiation, resulting in limited performance for solar installations. A second characteristic of the solar resource is its intermittent nature, which is due to the daily (day/night) and seasonal cycle, but also to meteorological effects such as the passage of clouds or heavy cloud cover. Thus, to meet these constraints, the catalysts developed must have photocatalytic properties that allow the use of the solar resource and manage the fluctuations linked to this resource. This implies catalysts with both photo-excitation characteristics to produce radicals and sorption characteristics to store the molecules during low sunlight phases.

Among the support materials envisaged, clay minerals are considered promising owing to their 76 interesting inherent multi-functional properties such as their adsorption capacity, high surface 77 area, multiscale porosity, and ability to be grafted by chemical compounds [18–25]. In this 78 79 respect, some authors more recently reported CeO₂_WO₃_Palygorskite/TiO₂ catalysts elaborated by impregnation method [26] and TiO₂ supported on three natural clays with 80 81 different morphologies carried out by sol-gel route [27]. In the same way, we reported some 82 years ago the immobilization of TiO₂ anatase nanoparticles (NPs) with an average size of 10 nm onto particles surfaces of beidellite [28] and nanofibers surfaces of palygorskite [29] via a 83 84 colloidal sol-gel route. Beidellite and palygorskite were both natural clay minerals sampled in Morocco from Agadir basin and Marrakech High Atlas regions, respectively. They were 85 purified, characterized, and specifically functionalized to be used as catalytic supports in the 86 wet synthesis methods employed [30,31]. Photocatalytic tests performed in presence of these 87 TiO₂ supported clay nanocomposites towards the degradation of Orange G dye (OG) were 88 promising in that they interestingly exhibited a photoactivity normalized to the TiO₂ amount 89 higher than that of pure TiO₂ micropowder commercially available as Degussa P25. 90

91 Nevertheless, the synthesis of these supported photocatalysts is generally achieved via colloidal 92 [21,29] or solvothermal [32] or pillaring [20] wet routes, which on one hand required previous 93 purifying and/or homoionisation with Na⁺ ions of clay minerals. On the other hand, these 94 elaboration methods involved several steps so that they are laborious, time and solvents 95 consuming and yielding to small laboratory amounts. Thereby, they have a low capacity to be 96 transferred to industrial scale. Thus, there is a need to find out alternative synthesis method to97 overcome these wet routes issues.

The aim of this work is to show the potential of these materials for a solar photo-oxidation 98 application. It is divided into three parts. The first part aims at elaborating TiO₂ supported 99 palygorskite by a free-solvent one-pot mechanochemical method starting from powders of raw 100 palygorskite (Pal) clay and titanyl sulfate (TiOSO₄·2H₂O) precursor. This dry synthesis route 101 102 was smartly based on *in-situ* reaction between carbonates accessorily available in raw clay and 103 a Ti molecular precursor to immobilize resulting TiO₂ particles on the surface of palygorskite fibers or in the vicinity of their surface in the tangle of fibers. Synthesis conditions were 104 105 optimized by studying the influence of the composition ratio palygorskite clay / TiOSO4 2H₂O and grinding velocity. The second step is to select the optimal conditions for the development 106 of a material in order to control the photocatalytic properties. The so elaborated TiO₂-Pal 107 108 nanocomposites were afterwards characterized at a multi-scale level using in situ X-ray diffraction (XRD (X'pert powders Philips diffractometer)), Scanning Electron Microscopy 109 (SEM (SEM-FEG, Hitachi S-4500)) and Transmission Electron Microscopy (TEM (JEOL JEM 110 111 2010)). The originality lies in a global approach starting from the elaboration of the catalysts, their optimization to test their photocatalytic potential. The aim being to evaluate the photo-112 113 oxidation capacities under solar conditions, our step-by-step approach involves three stages. Measurements are carried out under controlled irradiation conditions, both in terms of spectral 114 distribution and flux density (artificial UV and solar irradiation) with the aim of modelling the 115 116 photo-oxidative capacities of the catalysts. Finally, the photo-oxidation capacities are measured and modelled under natural irradiation, taking into account the characteristics of solar radiation: 117 spectral distribution and flux density discontinuities. Thus, photo-activities were evaluated 118 towards the removal from aqueous medium of Orange dye (OG) considered as a model pollutant 119

- under UV radiations and solar light by using indoor lab-scale reactors equipped with an artificial
- solar light and a larger outdoor pilot irradiated by natural solar light.

122 **2. Experimental details**

123 **2.1- Materials and synthesis**

Natural raw clay, involved herein, was sampled from Marrakesh – High Atlas region. Detailed 124 characterizations performed by Rhouta et al. [30] on this clay revealed that it is made up of fine 125 clay fraction (< 2 μ m) (> 65%) along with accessory minerals, namely quartz (< 5%) and 126 carbonates (≈ 30 wt%) in the forms of calcite (Mg_{0.03}Ca_{0.97}CO₃) and ankerite 127 $(Ca_{1.01}Mg_{0.45}Fe_{0.54}(CO_3)_2)$. The extracted fine fraction (< 2 µm) was found to be exclusively 128 composed of fibrous clay minerals, namely palygorskite (≈ 95%) exhibiting predominant 129 dioctahedral character and deficiency into zeolitic water along with sepiolite ($\approx 5\%$). The 130 131 composition of this palygorskite was found on the basis of 26 oxygens to be $(Si_{7.97}Al_{0.03})(Mg_{2.17}Al_{1.46}Fe_{0.40}Ti_{0.05})(Ca_{0.03}Na_{0.07}K_{0.03})O_{20.18}(OH)_{1.94}(H_2O)_{3.88}, 2.43 H_2O.$ Its 132 Cation Exchange Capacity (CEC), Brunaeur-Emmet-Teller method (BET) specific surface area 133 and total porous volume were assessed to be 21.2 meq.100 g⁻¹, 116 m².g⁻¹ and 0.458 cm³.g⁻¹ 134 respectively. 135

The Ti precursor considered herein is the titanyl sulfate (TiOSO₄·2H₂O) purchased from Aldrich
and used without further treatment.

In contrast to our previous works [24-27, 30] in which the synthesis of clay minerals supported photocatalysts implied several laborious and time-consuming steps including purifying, homoionisation and fine fraction recovering before its functionalization, the elaboration herein was accomplished according to a one-pot route starting directly from raw clay. Indeed, given amounts of raw palygorskite clay and TiOSO₄·2H₂O powders were introduced in an alumina jar of planetary ball mill PM100 from Retsch for being grinded using six alumina balls of 20 mm in diameter. The synthesis of TiO₂ supported on palygorskite fibers was achieved by *in*- 145 situ reacting carbonates accessorily present (≈ 30 wt%) in raw clay with TiOSO₄·2H₂O precursor according to different conditions. Indeed, the amount of each reactant was chosen in 146 147 such a way to have a total mass of reactants of 50 g per run by considering two raw clay / TiOSO₄ mass ratios of 0.6/1 and 3.3 /1, which corresponded to mixtures in which both 148 compounds were in stoichiometric or in excess proportions, respectively. The grinding was 149 150 performed at a rate of 500 rpm for different times (10, 15, 20 and 30 min). As-grinded samples 151 were thereafter air annealed for 1 h at different temperatures (400, 500, 600 and 700°C). Samples were designated Pal_TiO₂(x)_ G_{y} _Tz by referring to the composite nature of the sample 152 153 (Pal_TiO₂), the reactant mass ratio ($x = TiO_2/Pal$), the grinding time (G_y) and the post-heattreatment temperature (T_z) . The obtained powdered materials were finally stored for further 154 155 uses.

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157 **2.2- Characterizations**

158 The structural changes of TiO_2 upon heating of as-grinded samples were analyzed in situ versus the temperature by XRD over the two-theta range 2_60 deg using a Bruker D8 159 Advance diffractometer equipped with a Vantec Super Speed detector and an Mid- Infrared 160 (MIR) radiation heating chamber (Bragg-Brentano configuration; Ni filtered Cu K_a radiation). 161 The diffractograms were recorded every 50°C from the room temperature to 950°C. An 162 isotherm was maintained at each level for 40 min to record the pattern then the temperature was 163 increased using a ramp of 1 deg.s⁻¹ up to the next level. Also XRD at room temperature was 164 recorded on TiO₂_Pal samples prior to air annealing at different temperatures in the same 165 angular range using a Seifert XRD 3000TT diffractometer equipped with a graphite 166 monochromator (Bragg-Brentano configuration; Cu K_{α} radiation). 167

168 The microstructure, morphology and uniformity of the samples were analyzed at 169 different steps of the synthesis by a Scanning Electron Microscope (SEM; Leo-435VP)

170	equipped with an Oxford energy dispersive spectrometer (EDS, (EDS, KEVEX Si(Li) assisted
171	software Brüker)). Also, a JEOL JEM 2010 TEM equipped with a Tracor EDS analyzer was
172	used for characterizing clay particles and performing local elemental composition.

174 **2.3-** Photocatalytic activity

175 The photocatalytic activity of different samples was evaluated using three kinds of set-176 ups: a laboratory-scale set-up, an indoor solar simulator and an outdoor solar pilot.

As far as the first set-up is concerned, the degradation reaction was performed in a batch quartz 177 reactor $(40 \times 20 \times 36 \text{ mm}^3)$ placed in a thermostated chamber (25°C) under the UV light of a 178 179 lamp (HPLN Philips 125 W) emitting at 365 nm. The reactor was irradiated with a photon flux of about 100 mW.cm⁻² by adjusting the distance to the lamp so that it simulates the UV intensity 180 of solar spectrum on the earth [33]. This lamp was chosen because the OG absorption is 181 182 negligible at this wavelength and, as a result, the direct photolysis of the solution (*i.e.* without photocatalyst) was found negligible for more than 24 h. The photoactivity of different samples 183 184 was assessed at pH around 6 by measuring the decomposition rate of OG in aqueous solutions 185 containing the supported photocatalyst according to a procedure previously reported [34] and adapted to the dispersion of supported powder photocatalysts in previous papers [28,35]. The 186 photocatalyst powder of TiO2_Pal nanocomposites was added to 25 cm³ of OG solution (10⁻⁵ 187 M) in an amount equal to 0.8 g.dm⁻³. This catalyst mass was found as optimum to avoid excess 188 of catalyst and to ensure an efficient absorption of photons [36]. The dispersion was agitated 189 190 with an inert Teflon magnetic stirrer. To determine the dye concentration, aliquots were taken from the mixture at regular time intervals and centrifuged at 12500 rpm for 5 min. The OG 191 concentration in the supernatant was determined by measuring the absorbance at 480 nm using 192 193 a UV-VIS-NIR spectrophotometer (Perking Elmer lambda 19).

Sunlight is particularly weak in the UV spectral region and, in addition, it changes along the 194 195 day due to day-night and seasonal cycles [37,38]. Thus, prior optical and kinetic studies have to be performed to determine on one hand optimal catalyst mass allowing the maximum (\approx 196 197 90%) of the absorption of photons (considered as reactant) and, on the other hand, kinetic constants to achieve process efficiency in an outdoor solar pilot [37,39]. Therefore, optical 198 properties of photocatalysts suspensions differing in mass concentrations were studied by an 199 200 optical experimental setup described in details elsewhere [37,39]. Practically, the light supplied by a source simulating solar radiation (maximum intensity equal to 1000 W.m⁻² in the full 201 spectral range, *i.e.* 50 W_{uv}.m⁻² in the UV range) came across a Polymethyl methacrylate 202 (PMMA treated to be UVA transparent, Sunactive® GS2458) cell 2 cm thick (i.e. optical 203 length) containing under stirring the TiO2_Pal dispersions whose concentrations ranged from 204 205 0.1 to 4 g.L⁻¹. This cell was placed at the inlet of an integrating sphere. The direct radiation as well as scattered radiation transmitted hence collected from the integration sphere were directed 206 207 via an optical fiber towards a spectrophotometer permitting the detection of wavelengths 208 ranging from 250 to 1100 nm (ultraviolet, visible and near-infrared). The transmission of light was hence assessed versus catalyst powder concentration to determine the mass concentration 209 210 at which the 90% absorbency was reached. Afterwards, kinetic study of OG dye degradation 211 was carried out versus mass concentration of catalysts powders dispersion by using an indoor solar simulator set-up. For a given mass concentration the catalyst dispersion was kept under 212 213 stirring in a 150 mL quartz bucker (suspension height of 5 cm), then it was irradiated by sunlight simulator (1000 W.m⁻²) under constant photons flow. At different time intervals, aliquots were 214 215 picked up from media, centrifuged and resulting supernatants analyzed by spectrophotometry to determine the dye concentration according to the way described above. 216

Afterwards, the efficiency of TiO_2 _Pal catalysts towards the photodegradation under sunlight of OG dye was tested by using an outdoor solar pilot described in details by Janin *et al.*, [40]

(picture given in the supplementary figure 1S). Briefly, a mass concentration of TiO₂ Pal that 219 permitting 90% of photons absorption was dispersed under stirring in an OG dye solution and 220 stored in a stainless steel tank (with a maximum volume of 50 L). Then 6 L of effluent were fed 221 into a flat solar panel $(30 \times 100 \times 2 \text{ cm}^3)$ from the bottom to the top via a liquid pump and 222 recirculated into the thank. This panel was covered with a special PMMA plate to prevent 223 evaporation of compounds and to ensure watertight. This flat reactor was oriented southwards 224 and inclined at an angle around 42° in order to maximize the average solar irradiation 225 throughout the year. This was controlled by two equipments: a pyranometer whose wavelength 226 227 domain varies between 380 to 2800 nm and an UV radiometer (UVA-Sensor CT-UVA 3) measuring the corresponding spectrum in the range of 310 - 400 nm; 228

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230 3. Results and discussion

231 **3.1- Structural characterizations**

Figure 1 shows an overview of the room temperature XRD diagrams of the two starting materials after a grinding for 30 min at 500 rpm: titanyl oxysulfate precursor (TiOSO₄) and palygorskite-rich raw clay (Pal). Also, the XRD diagrams of a mixture of TiOSO₄ and Pal clay in a mass ratio of 0.6/1 after grinding for 30 min at 500 rpm is reported (Pal_TiOSO₄).



Fig.1: Room temperature XRD patterns of TiOSO₄, raw clay (Pal) and Pal_TiOSO₄ mixture in
mass ratio of 0.6/1 after grinding of each of these samples at 500 rpm for 30 minutes.

The diffractogram pertaining to the milled titanium precursor revealed only the reflections at 240 the 20 angular positions of 10.67° (8.35 Å), 10.95° (8.20 Å), 18.61° (4.79 Å), 19.07° (4.69 Å), 241 21.43° (4.16 Å), 25.44° (3.50 Å), 25.70° (3.47 Å), 28.47° (3.13 Å), 29.25° (3.05 Å), 37.70 (2.38 242 Å), 38.61 (2.33 Å) corresponding to the compound TiOSO₄ (ICDD file N°14-503) without 243 evidence for the formation of any other crystalline phase, especially TiO₂. In the same sense, 244 245 the XRD pattern of the grinded Pal raw clay showed several reflections corresponding to the fibrous mineral, *i.e.* palygorskite, at the angular positions described by Rhouta *et al.*, [30]. The 246 characteristic basal reflection is at $2\theta \approx 8.38^{\circ}$ (10.56 Å) in addition to the peaks ascribed to the 247 accessory minerals, namely the carbonates in the form of calcite (C) Mg_{0.03}Ca_{0.97}CO₃ (ICDD 248

file: 01-089-1304) and ankerite (A) Ca_{1.01}Mg_{0.45}Fe_{0.54}(CO₃)₂ (ICDD: 01-084-20 2066) as well
as quartz (Q) (ICDD file: 03-065-0466) whose strongest reflections were observed at 29.46°
(3.03 Å), 30.82° (2.90 Å) and 26.64° (3.34 Å), respectively.

252 This result showed that both the palygorskite and the accessory phases, constituting the raw Pal clay, were apparently not affected by the grinding operation in the mechanosynthesis jar since 253 254 their crystallographic structure remained stable. On the other hand, these grinding conditions 255 of the Pal_TiOSO₄ mixture caused a drastic decrease in the XRD intensities of TiOSO₄ reflexions at the point to disappear, as well as those of the carbonates (calcite and ankerite), 256 while preserving the clay structure confirmed by the persistence of the corresponding peaks. 257 258 The disappearance of TiOSO₄ reflexions can be due to its amorphisation but the major event during grinding is the trigger of reaction (1), according to which TiOSO₄ reacts with the 259 carbonates present among the accessory minerals in the raw clay Pal: 260

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$$\text{TiOSO}_4 \cdot 2\text{H}_2\text{O} + \text{Pal}/(\text{Na}_2, \text{Ca}, \text{Mg})\text{CO}_3 \rightarrow \text{TiO}_2/\text{Pal} + (\text{Na}_2, \text{Ca}, \text{Mg})\text{SO}_4, 2\text{H}_2\text{O} + \text{CO}_2\uparrow$$
 (1)

262

The occurrence of this reaction was further confirmed by the production of the hydrated sulfate compound CaSO₄·0.5H₂O evidenced by the presence in the composite sample Pal_TiOSO₄ of the characteristic peaks noted S on the diffractogram (Fig. 1) at angular positions 2 θ of 14.71° (6.00 nm), 29.71° (3.00 nm) and 31.86° (2.80 nm) in agreement with the ICDD file 01-081-1848. The near-consumption of the TiOSO₄ precursor revealed by XRD may be due to the fact that the low value of the mass ratio TiOSO₄:clay (0.6:1) would be the driving force of reaction (1).

Furthermore, the reaction (1) also showed that titanium dioxide has to be formed but no diffraction peaks of this oxide was observed in Figure 1 suggesting that the as-prepared TiO_2 phase must be amorphous after the mechanosynthesis step (TiOSO₄:clay = 0.6:1; 500 rpm; 30 min). Subsequently, a mixture of TiOSO₄ and Pal raw clay with the mass ratio 3.3/1, *i.e.* with
Ti precursor in excess was grinded for 30 min at 500 rpm. Then, XRD patterns of this composite
sample were recorded *in-situ* in air as a function of the temperature (Figure 2).



Fig. 2: X-ray diffractograms recorded *in-situ* in air as a function of the sample temperature of a TiOSO₄/Pal (3.3/1) mixture after mechanosynthesis (Step 1). These diagrams were recorded every 50°C; they represented the effect of calcination as 2nd step of the synthesis process. They showed the main compounds that disappeared were TiOSO₄ (\bigstar), calcite (C), ankerite (A), while the compounds that formed were CaSO₄·xH₂O (\blacktriangledown), CaSO₄ (\bullet), TiO₂ anatase (\blacklozenge), and the most thermally stable accessory compounds like quartz (Q).

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For as-milled sample, the diagram recorded at 50° C revealed that, as mentioned above, amorphous TiO₂ and the hydrated CaSO₄ product were formed during the grinding stage as a result of the above reaction (1). In this respect, a significant proportion of the carbonates were consumed by the reaction with TiOSO₄ (that was in excess) has shown by the decrease of their XRD intensity. This was further supported by a huge decrease in carbonate content determined by coulometric analysis. Indeed, the carbonate content has been reduced approximately by a factor of 2, from 28 ± 2 wt % in the raw clay Pal to $13 \pm 2\%$ wt % in the grinded TiOSO₄ / Pal (3.3/1) mixture.

292 By increasing the temperature, it can be seen that the crystallization of anatase TiO₂ started from 450°C as shown by the emergence of the corresponding peaks at angular positions 20 of 293 about 25.44° (3.51 Å) and 38.53° (2.33Å) (ICDD Data Sheet 01-021-1272). As the temperature 294 increased further, the anatase peaks became more intense and well resolved, indicating the 295 296 improvement of its crystallization. The anatase peaks remained observable up to 700°C, without the appearance of any XRD peaks of the less photoactive rutile variety. This strongly denoted 297 the remarkable stability of the anatase formed under the conditions of this study. Palygorskite 298 299 peaks were observable up to 400°C, beyond which their intensities severely reduced. This can be ascribed to the occurrence of structural collapse of the clay mineral. However, this can be 300 also due to an absorption filter induced by the conformal coverage of anatase particles on 301 palygorskite fibers as suggested by Rhouta et al., [28] and Bouna et al., [29] and confirmed by 302 303 the TEM analyses shown in the next section.

In order to examine the effect of grinding time on the reactivity of TiOSO₄ with the carbonates present in the Pal raw clay, a mixture of the two compounds in a ratio of 0.6 / 1 was ground at 500 rpm for 10, 15, 20, 25 and 30 min and then air annealed at 600°C for 1 h. The superimposition of the corresponding XRD diagrams (Fig. 3) clearly showed the formation of anatase TiO₂ for all the samples investigated regardless grinding time. Nevertheless, the anatase crystallinity appeared to improve with the grinding time as evidenced by the gradual narrowing of it characteristic peak at about 20 around 25°. In this respect, by referring to quartz, which is an insensible and stable accessory mineral under the applied grinding conditions, the increase
of the intensity ratio of the main anatase/quartz peaks is noteworthy. This probably reflected an
increase in the amount of anatase phase formed with the increase of grinding time.

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Fig. 3: Stacking of XRD patterns of TiO₂_Pal photocatalysts prepared by grinding TiOSO₄ and
Pal raw clay in a mass ratio of 0.6/1 during 10, 15, 20, 25 and 30 min at 500 rpm.

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319 **3.2- Microstructural characterizations**

320 XRD results were further supported by TEM characterizations. Indeed, depending on the 321 grinding conditions, TEM analysis after the first step showed palygorskite fibers of 30-70 nm 322 in diameter and 650 nm in length mixed with TiOSO₄ platelets of 400-600 nm with an average 323 thickness of 3-5 μ m. At this stage, no evidence for crystallized anatase nanoparticles (NPs) was 324 found confirming this oxide phase was amorphous. After annealing at 600°C in air for 1 h, TEM 325 revealed that palygorskite fibers were entirely wrapped with homogeneous monodisperse and spherical TiO₂ anatase NPs 10-15 nm in diameter (Fig. 4a). Only a few palygorskite fibers were
not covered as revealed by EDS analysis. For most of them, a granular and conformal TiO₂
coverage formed a uniform cladding (Fig. 4b).

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Fig. 4: TEM micrographs and corresponding electron diffraction pattern indexed as TiO₂ anatase (a) and EDS local analysis showing uncoated (point 2) and uniformly coated point 1 and 3) palygorskite fibers (b).

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336 **3.3-** Photocatalytic properties under UV light

337 The photocatalytic activity of different TiO₂-Pal composite samples elaborated by mechanosynthesis was tested towards the degradation under UV light of OG dye by using the 338 laboratory set-up described in the experimental section (2.3). The results in Figure 5 showed 339 that photocatalytic efficiency significantly increased with the amount of TiOSO4 used in 340 photocatalyst synthesis and thus the amount of TiO₂ present in TiO₂_Pal nanocomposite. This 341 342 was evidenced by on one hand the strong decrease of OG concentration for instance in the first 10 min, revealing a high initial degradation rate of the pollutant around $2.6 \times 10^{-2} \text{ s}^{-1}$ for TiO₂/Pal 343 mass ratio of 3.3/1 with respect to that determined for the mass ratio of 0.6/1 ($\approx 1.5 \times 10^{-2} \text{ s}^{-1}$). 344 On the other hand, the almost total elimination of the OG dye from the solution was achieved 345 within about 2 h when using TiO₂/Pal mass ratio of 3.3/1 whilst almost 10% of the dye remained 346

347 after the same time of UV irradiation in presence of TiO_2/Pal mass ratio of 0.6/1 (but a plateau





Fig. 5: Change in the concentration of OG (where C_0 and C denote the concentrations of OG at the initial time t = 0, and after an elapsed time t of the photocatalytic reaction) vs. UV irradiation time of the samples, differing in TiOSO₄/ Pal mass ratio (0.6/1 and 3.3/1), grinded during 30 min at 500 rpm and afterwards air annealed at 600°C for 1 h.

The influence of annealing temperature on the photocatalytic activity towards the removal of OG dye of TiO₂_Pal was afterwards studied on the most active one corresponding to TiO₂/Pal mass ratio = 3.3/1. Figure 6 interestingly showed that the unheated as-milled Pal_TiOSO₄ mixture (sample Pal_TiO₂_(3.3)_30) already exhibited photoactivity under UV irradiation. Indeed, 50% of the initial quantity of OG decreased after 2 hours of irradiation. This is due to the formation of photoactive TiO₂ resulting, as beforehand mentioned, from the heterogeneous

solid-state reaction between TiOSO₄ and the carbonate phases (calcite + ankerite) present as accessory minerals in the Pal raw clay (reaction 1). Although the TiO₂ formed in as-grinded material was amorphous, as ascertained above by XRD, this is not exceptional since it was already shown that amorphous TiO₂ can be active in photocatalysis under UV irradiation [41]. Indeed, from a general point of view, even if amorphous semiconductors have less efficient electronic properties than their crystallized counterparts, they nevertheless find specific applications as in photovoltaic.

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Fig. 6: Change in the concentration of OG vs. UV irradiation time of TiO_2 _Pal elaborated by mechanosynthesis (TiOSO₄/Pal mass ratio = 3.3/1, grinding at 500 rpm for 30 min) and air annealing at 400, 500, 600 and 700°C.

Samples annealed at 400 and 500°C showed quite similar photo-degradation curves of the OG 374 and paradoxically showed a lower photoactivity compared to as-grinded sample, allowing the 375 removal of only 15% after 2 hours of illumination. This annealing temperature range is 376 377 therefore insufficient to ensure a good crystallization of TiO₂ in the form of anatase but, on the other hand, it seems to be sufficient for inducing a detrimental diffusion of species that trap 378 excitons (Pal clay cations for example), thus affecting the photocatalytic efficiency of the 379 composite. Calcination at 600°C promoted the most efficient active palygorskite-supported 380 TiO₂ photocatalyst since it allowed almost complete degradation of the OG after 2 hours of UV 381 irradiation. This improvement was attributed to the good crystallinity of TiO₂ anatase upon the 382 383 sample calcination at 600°C as confirmed above by XRD. On the other hand, when the sample was annealed at 700°C, photocatalytic activity again decreased. This reduction in photocatalytic 384 activity may be due to the growth of TiO₂ particles beyond the optimal nanocrystalline size, 385 386 which promotes the recombination of charge carriers. Also, a partial phase transformation of the more active anatase TiO₂ phase into the less photoactive rutile variety as reported by Bouna 387 388 et al., [29] and Rhouta et al., [28] is not excluded although none of the corresponding peaks 389 have yet been detected at 700°C, the presence of other crystalline phases can hide them.

390

391 3.4- Photocatalytic properties under sunlight

392 Owing to its high photoactivity under UV evidenced above by the tests on the laboratory set-393 up, the TiO_2 _Pal nanocomposite with mass ratio TiO_2 /Pal of 3.3/1 was thereafter retained for 394 performing the study of photocatalytic activity under sunlight towards the removal of OG dye. 395

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- 397
- 398

399 3.4.1- Activity under constant solar irradiation – indoor experiments

For studying optical properties of TiO_2 -Pal, the change in transmittance (%), defined as the ratio of transmitted (I_t) to incident (I_i) intensities, was followed versus catalyst slurry mass concentration. Figure 7 showed that the necessary amount for the absorption of 90% of incident



404 Fig. 7: Variation of transmittance in the spectral range 250 -1100 nm (artificial solar light
405 1000 W.m⁻²) as a function of the catalyst dispersion concentration in an aqueous suspension.
406

407 light simulating sunlight (1000 W/m²) for a catalyst dispersion thickness of 2 cm (C_{2cm}) was 408 around 2.36 g/L, *i.e.* 1.34 g/L of TiO₂ taking into account its proportion of 57 wt% assessed by 409 Individual Counting by single Particle (ICP) in the nanocomposite. Photocatalytic tests in batch 410 mode using the indoor solar simulator described in the experimental section were performed by 411 varying the amount of TiO₂_Pal dispersed in OG solution of 5 cm of thickness around the 412 optimum amount (C_{5cm} = $2/5 \times C_{2cm} \approx 0.95$ g/L) allowing maximum absorption.

Figure 8 showed that very low amount of TiO₂ Pal, e.g. 0.25 g/L, exhibited a relatively low 414 photocatalytic activity leading to the removal of about only 40% of OG after 150 min of 415 irradiation with an initial photo-degradation rate of about 5×10^{-3} s⁻¹. Nevertheless, by increasing 416 417 the TiO₂_Pal mass concentration beyond 0.5 g/L up to 1.5 g/L, a better efficiency occurs and a quite similar photocatalytic behavior of the photocatalyst towards OG elimination was 418 observed. Indeed, the OG photodegradation considerably increased as evidenced on one hand, 419 420 by the strong decrease of OG concentration for instance in the first 20 min, revealing a high initial degradation rate ($\approx 7 \times 10^{-3} \text{ s}^{-1}$) of the pollutant and, on the other hand, by the elimination 421 of almost 60% of OG dye from the solution within the same time of irradiation. These results 422 showed that around the optimal mass concentration of TiO₂_Pal (≈ 0.95 g/L) beforehand 423 determined by optical measurements, the catalyst exhibited maximal photoactivity for dye 424 removal due to highest light absorption. Thus, an increase in the mass concentration of the 425 photocatalysts beyond its optimal mass concentration is unnecessary since the maximum 426 absorption of incident light has already been reached. This is in agreement with Plantard et al., 427 [42] who reported that for a given pollutant and a fixed initial concentration, the degradation 428 429 kinetic constants obtained for the same photocatalyst depended solely on the quantity of 430 photons absorbed. As a result, degradation by photocatalysis is function to the concentration of the catalyst in suspension, up to a threshold value at which all the incident radiation is almost 431 absorbed. 432

433

434



435

Fig. 8: Variation of the relative OG concentration versus time (where C_0 and C denote the OG concentrations at the initial time t = 0, and after an elapsed time t) in the 1000 W.m⁻² solar simulator recorded for different mass concentrations of the TiO₂_Pal composite photocatalyst.

Besides photocatalyst mass concentration in the aqueous suspension, photons fluxes also constituted a crucial parameter influencing heterogeneous photocatalysis process especially under solar light characterized by the variation of irradiation intensity. Figure 9, representing OG proportion change versus solar irradiation time at different photons fluxes in presence of 0.94 g/L of TiO₂_Pal revealed that in general, and as expected, an increase of the radiation intensity led to an increase of the OG pollutant degradation. This is the result of the increase of quantity of photons absorbed. It is obviously well known that this is only the UV part of the 448 solar spectrum that is effective for photocatalytic treatment with TiO₂ as active component. 449 With the solar simulator UV intensity is directly proportional to the total intensity delivered by 450 the lamp. It is almost the case for solar natural irradiation (even if the UV percentage is 451 influenced by the weather conditions). So, we made the choice to keep the total radiation 452 intensity received as the reference. It is the energy flux received by the surface of a reactor 453 working in outdoor conditions.

454

For being transposed to solar photocatalysis treatment, OG photodegradation kinetic should be modulated taking into account the irradiation intensity key parameter. In photodegradation of a target molecule (OG herein) using a slurry catalyst (TiO₂_Pal), rates can be approached by the modified Langmuir-Hinshelwood (L-H) relationship (1) [43]:

459

460
$$\frac{dC}{dt} = \mathbf{k}(\mathbf{I})\frac{C}{\mathbf{1}+\beta C} = -(\alpha \cdot \frac{S}{V_{T}} \cdot \mathbf{I})\frac{C}{\mathbf{1}+\beta C}$$
(1)

With k(I) the apparent kinetic constant (s⁻¹) function of the intensity, C the OG concentration (g.L⁻¹), β (L.g⁻¹) a constant determined by comparison between calculated and experimental data. In the simplest case, k(I) is directly proportional to the intensity I received by unit volume of solution (V_T in L) with an irradiated surface S (m²). α (J.L⁻¹) is the second unknown parameter. If necessary, a dependence of k according to a law in power of the irradiation can be introduced to account of additional limiting phenomena such as the charge recombination.



467

468 Fig. 9: Variation of the OG proportion versus time (s) upon irradiation simulating sunlight at
469 400, 600 and 1000 W.m⁻² in presence of TiO₂_Pal (0.94 g/L) composite photocatalyst.

- 470
- 471

472 For a constant irradiation intensity, the integration of equation (1) leads to the expression (2):

473
$$f\left(\frac{c}{c_0},\beta\right) = \ln\left(\frac{c}{c_0}\right) + \beta \cdot C_0 \cdot \left(\frac{c}{c_0} - 1\right) = -k(I) \cdot t$$
(2)

To be valuable, the value of β has to be constant whatever is the irradiation. It is determined simply by considering the left part of equation 2 depends linearly on time. Figure 10 depicted experimental profiles of the logarithmic adimensional concentration C/C₀ as a function of the time with a value of $\beta = 0.65$ L.g⁻¹. It showed linear dependence with a satisfactory agreement. This ascertained pseudo first-order kinetics whose the slope k depends on the intensity of the irradiation according to the relationship (3).

480
$$k(I) = \frac{s}{v_T}. \alpha. I \tag{3}$$



Fig.10: Linear variation of $Ln(C/C_0)$ of OG versus time (s) for different irradiation intensities simulating sunlight in the presence of 0.94 g/L of TiO₂_Pal catalyst for $\beta = 0.65$ L.g⁻¹ confirming a pseudo first order kinetics. Continuous colored lines represent the best agreement between the experimental data and a linear regression for each level of irradiation.

Likewise, the variation of the apparent kinetic constant k versus irradiation intensity (I) (Fig.
11) showed a linear trend supported by a reasonable degree of accuracy of about 0.99. Hence,
the value of the kinetic constant (α), deduced from the slope, can be assessed to be around 1.5
L/kJ.

481



494 Fig. 11: Linear variation of the kinetic constant k of OG photodegradation versus the
495 irradiation intensity I simulating sunlight onto TiO₂_Pal photocatalyst slurry. The continuous
496 gray line simulates the data according to equation (3).

493

By introducing the values of α and β in the expression (2), the C/C₀ plots versus time show a very good fitting between the experimental and simulated data for all irradiation flux densities, as expected from the good agreement already reported in Figures 10 and 11. This result validates the proposed kinetic model described by equation (2).

502

503 3.4.2- Activity under natural discontinuous irradiation - outdoor solar pilot

The photocatalytic tests for the removal of OG dye from aqueous solution were carried out in presence of TiO_2 _Pal slurry by using an outdoor solar pilot and according to conditions described in the experimental section (2.3). Nonetheless, owing to the variation of solar

irradiation along a day, the change of OG dye during these tests should be versus the cumulative energy Q_g (J.L⁻¹) received by the reactor given by the relationship (4) [44]:

509
$$Q_g = \frac{S}{V_T} \cdot \int_0^{t_i} I(t) dt$$
 (4)

with t_i (s) the time at which an aliquot was taken for being centrifuged and analyzed by the spectrophotometer and I(t) the solar irradiation intensity measured over time using the sensors mounted on the pilot as described in the experimental section.

513

Figure 12 shows, like data obtained with the indoor solar simulator in "batch" mode, a decrease in OG dye proportion upon solar irradiation, which thus confirms the photoactivity of TiO₂_Pal under solar light. In this respect, 50% of the OG dye was removed with a cumulated energy of 74.2 kJ/L whilst the total removal of the dye required a cumulated energy of 250 kJ/L.



518

Fig. 12: Change of the experimental relative concentration of the OG (×) and simulated data
(dotted curve) versus the total cumulated energy obtained for TiO₂_Pal composite

521 photocatalyst.

Besides, by introducing the numerical values of α (1.5 L/kJ) and β (0.65 L.g⁻¹) beforehand determined in the relationship (5) derived from the above one (2) after substituting I(t) by Q_g(t), a good agreement between simulated and experimental data was observed with a good degree of accuracy.

527
$$\frac{C(t)}{C_0} = e^{-[(\alpha \ \frac{s}{V_T} Q_g(t)) + (\beta C_0(\frac{c}{C_0} - 1)]}$$
(5)

This confirmed the validation of the kinetic model proposed and ascertained that the OG photodegradation in presence of TiO_2 _Pal slurry was of first order with a kinetic constant of **1.5** L/kJ and did not involve adsorption phenomenon.

531

532 4. Conclusion

TiO₂ photocatalyst supported onto natural palygorskite clay mineral fibers (TiO₂_Pal) was successfully elaborated by a one-pot dry mechanical method. This elaboration process is very interesting because it directly uses palygorskite clay without the multiple purification steps that wet processes frequently required. Furthermore, it is solvent free, eco-friendly and based on smart *in situ* solid-state reaction between carbonates accessory minerals contained in the raw clay and a cheap molecular Ti precursor. As a result, it can be easily scaled up to an economically viable large-scale production.

Structural analysis and microstructural characterization using XRD and TEM respectively showed that TiO_2 _Pal, revealed homogeneous monodisperse TiO_2 NPs in the form of anatase wrapping palygorskite fibers. This sample was shown to be the most photoactive under UV irradiation towards the removal of OG dye from aqueous solution upon the use of a laboratory set-up. Then photocatalytic tests carried out by an indoor solar simulator set-up on OG dye solution containing an optimal mass concentration of TiO_2 _Pal nanocomposite was found to

546	assure the maximum of the elimination of OG. In this respect, a good agreement was
547	demonstrated between experimental and simulated data using a pseudo first order kinetic model
548	for the dye photodegradation. The same photocatalytic behavior was also obtained with an
549	outdoor solar pilot confirming the photoactivity of these supported TiO2_Pal photocatalysts
550	under sunlight .
551	
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