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Joanna Urban, Abdelaziz Jouaiti, Nathalie Gruber, Géraud Delport, Gaëlle Trippé-Allard, et al.. Using chiral ammonium cations to modulate the structure of 1D hybrid lead bromide perovskites for linearly polarized broadband light emission at room temperature. *Journal of Materials Chemistry C*, 2022, 10 (34), pp.12436-12443. 10.1039/D2TC02040H . hal-03812431

HAL Id: hal-03812431

<https://cnrs.hal.science/hal-03812431v1>

Submitted on 12 Oct 2022

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Using chiral ammonium cations to modulate the structure of 1D hybrid lead bromide perovskites for linearly polarized broadband light emission at room temperature

Joanna M. Urban^{a‡}, Abdelaziz Jouaiti^{b*}, Nathalie Gruber^b, Géraud Delport^c, Gaëlle Trippé-Allard^a, Jean-François Guillemoles^c, Emmanuelle Deleporte^a, Sylvie Ferlay^{b*} and Damien Garrot^{d*}

We report on a series of low dimensional 1D enantiomerically pure and racemic lead bromide perovskite compounds of the formula **BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ (**BuA**, cation derived from 2-Butylamine, **MBA** from Methylbenzylamine and **EBA** from α - Ethylbenzylamine) incorporating monovalent chiral ammonium derivative cations. A series of 6 enantiomerically pure compounds is described. The compounds are fully characterized from a structural point of view using X-Ray diffraction on single crystals and XRPD. Their photoluminescence emission properties have been carefully investigated. The use of cations of different sizes allows tuning the distances between the inorganic one-dimensional (1D) chains. All compounds show broadband below-bandgap emission, attributed to Self-Trapped Exciton (STE) recombination. The photoluminescence is characterized by a significant degree of linear polarization at room temperature, which can be explained by the structural anisotropy of the transition dipole moments and quantum confinement of the excitons in the 1D material. Our results indicate that the described compounds are of high interest for linearly-polarized broadband light-emitting devices and demonstrate the possibility of engineering the crystalline structure by the appropriate choice of the organic cation.

Introduction

Hybrid organic–inorganic perovskites (HOIPs) are a class of crystalline semiconductor materials displaying tunable structures of varying dimensionality.^{1,2,3,4,5} Their remarkable opto-electronic properties, including large absorption coefficients⁶ and long carrier diffusion lengths,⁷ make HOIPs promising candidates for applications requiring high quantum efficiency, most notably as light-harvesting materials for photovoltaics^{1,2,6} as well as active materials in light-emitting devices and photodetectors.⁸ The HOIP lattice is built from large organic cations and an inorganic network (large metal cations (Pb²⁺, Bi³⁺, ..) and halide anions). The structural features and photophysical properties of the material can be tailored by varying the ionic composition.⁹ Generally, introducing large organic cations which violate the Goldschmidt tolerance factor into the lattice¹⁰ disrupts the 3D structure, leading to the formation of lower dimensional 1D or 2D perovskite materials with optoelectronic characteristics different from their 3D analogs.

While the majority of reported lead halide perovskite belongs either to the layered 2D or 3D class,⁴ interest in 1D halide perovskites^{11,12} has recently been growing due to their inherent one-dimensional character, which can enhance charge carrier transport along a particular direction¹³, as well as their typically high crystallinity and consequently high quantum efficiencies and long carrier diffusion lengths,¹¹ which are of prime importance for the performance of perovskite-based nanoscale optoelectronic and photonic devices.^{14,15,16,17} The structural anisotropy of perovskite structures with lowered

dimensionality manifests in their electronic and optical properties and makes them highly attractive for polarization-sensitive optoelectronic devices.¹⁸ In addition, broadband emission in the visible range with high potential for white light emission applications has been recently observed for low-dimensional perovskite materials.^{3,19,20,21,28} It is generally believed that exciton self-trapping (Self-Trapped Excitons, STE) in the strongly distorted lattice lies at the origin of such broadband photoluminescence.^{3,19,22,23} The chromaticity of the emission can be tuned, for example, by varying the nature of the halogen,²⁴ making low-dimensional HOIPs promising white light emitters.

In lower-dimensional perovskite structures, optoelectronic properties can be controlled by the choice of cations with specific functionality. For example, chiroptical properties can be transferred to the perovskite structure by the intercalation of chiral cationic organic compounds.^{18,25,26,27,28,29} In the last five years, the interest in chiral Hybrid Organic-Inorganic Perovskites (c-HOIP) has grown, mainly due to their intriguing physical properties including nonlinear optical responses, circular dichroism, ferroelectricity, circularly polarized light absorption and emission and spin dependent transport.^{30,31,32}

1D c-HOIPs perovskites combine the advantages of chiral materials and low-dimensional lead halide perovskite structures and hold promise for a wide range of applications in life science, material science and in next-generation optical and spintronic devices.^{26,30}

In this work, we report the synthesis of one-dimensional compounds incorporating chiral (and for some of them racemic) cations derived from 2-Butylamine, **BuA** (cation **BuAH**⁺, in the following called **BuA**), Methylbenzylamine, **MBA** and α -Ethylbenzylamine, **EBA**, represented in figure 1. The choice of these cations is guided by their availability and ability to form

1D structures with lead perovskites. After a careful structural study series of 6 enantiomerically pure compounds and for two of them the racemic compounds, we characterize their photoluminescence properties and observe a broadband yellow to orange emission with a significant degree of linear polarization.

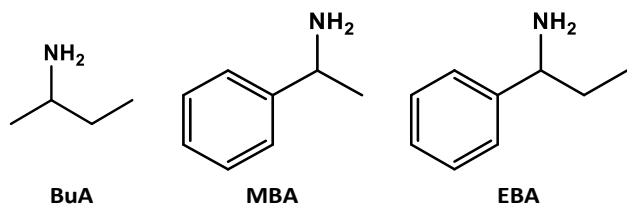


Figure 1: The amines used for the formation of lead-bromide species **BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃.

Results and discussion

Structural characterization

Three pairs of enantiomeric cations have been used in the following study. **BuA**, **MBA** and **EBA** are monovalent cationic species (see figure 1). To the best of our knowledge, lead-halide structures with **BuA** or **EBA** cations have not been reported previously. Concerning the **MBA** amine, several 1D structures have been already reported, among them the structure consisting of *R*-**MBA** combined with a PbCl₃, PbBr₃³³ or Pbl₃ inorganic sublattice and *R*-**MBA** and *S*-**MBA**^{25,34} as well as the racemic mixture of the cations combined with the Pbl₃ sublattice.¹⁴ Other related 2D structures involving mixtures of Pbl₄ and PbBr₄ have also been reported.²⁸

The enantiomeric compounds (*S*) and (*R*)-of **BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ are all isomorphous, crystallizing in the *P*₂₁₂₁₂₁ space group (see crystallographic table S1, ESI), belonging to non-centrosymmetric Sohncke space groups. The crystal structure is based on cationic protonated monovalent cations and the anionic inorganic PbBr₃⁻ chains formed by face-sharing metal-halide octahedra. The compounds have one-dimensional character on the molecular level, in contrast to materials with 3D stoichiometry where the low-dimensional character is due to purely morphological confinement.³⁵ For the six reported compounds, the inorganic chains are distorted (see figure 2) and the metal cations are in a distorted octahedral environment as shown by angles and distances reported in table S2, SI, and as demonstrated by the Shape analysis detailed below. The inorganic units are not helicoidal and thus not intrinsically chiral.

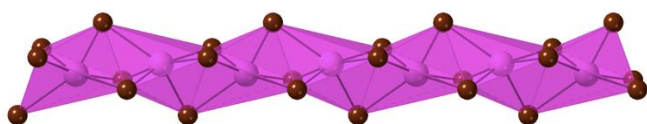


Figure 2: Schematic polyhedral representation of one PbBr₃⁻ anionic inorganic chain in the **BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ compounds.

In the unit cell, the 1D bromoplumbate arrays are separated by the chiral amines, as shown in figure 3, with an alternate disposition of the protonated amino groups. Within the unit cells, there are some hydrogen bonds formed between the amino groups and the Br atoms, with the N...Br distances in the 3.3940(13)–3.6653(13)Å range, as shown in table S2, SI.

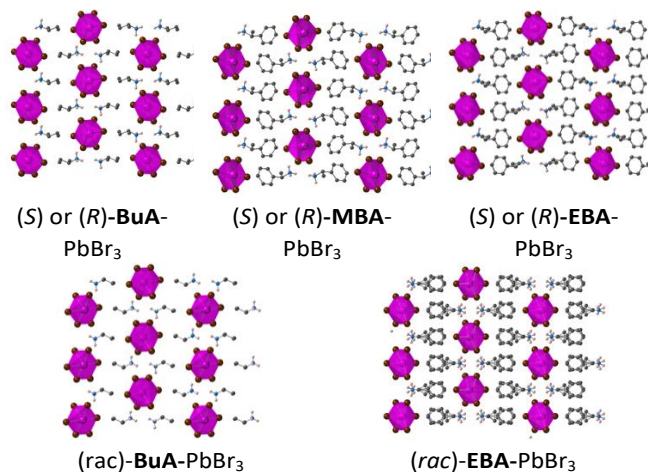


Figure 3: Representation of the structures of either (*S*)- or (*R*)-**BuA**-PbBr₃, **MBA**-PbBr₃, **EBA**-PbBr₃, (*rac*)-**BuA**-PbBr₃ and (*rac*)-**EBA**-PbBr₃, projection in the *xOz* planes. Only H atoms located on N are presented. For bond and distances see text.

1D compounds including cyclohexane (rather than phenyl) derivative amines have been reported not only for the *R* and *S* enantiomers but also for their racemic mixtures.³⁶ For this reason, the synthesis and characterization of the racemic analogs of the three enantiomeric compounds have been undertaken. The structures of the racemic (*rac*)-**BuA**-PbBr₃ and (*rac*)-**EBA**-PbBr₃ compounds are analogous to the ones reported for the corresponding enantiomers. (*rac*)-**MBA**-PbBr₃ reproducibly crystallizes into another phase, including solvent molecules. Despite many attempts we were not able to obtain an achiral phase structurally analogous to the ones observed for (*S*)-**MBA**-PbBr₃ or (*R*)-**MBA**-PbBr₃. The formation of a different phase for a racemic mix of the organic large cation molecules was previously reported for other lead halogen perovskites with chiral amino based cations and explained by a different packing of the molecules within the structure.^{3,37}

For the **BuA** and **EBA** cations, the corresponding racemic compounds crystallize in centrosymmetric space groups (monoclinic *P*₂₁/*n* for (*rac*)-**BuA**-PbBr₃, and orthorhombic *Pnma* for (*rac*)-**EBA**-PbBr₃). For both compounds, the PbBr₃⁻ anionic chains (analog to the ones of the enantiomerically pure compounds, represented in figure 2) and the ammonium cations are present in the unit cell, without solvent molecules (see figure 3). For (*rac*)-**EBA**-PbBr₃, the organic cations appear to be disordered. Bonds and distances are reported in table S2. As in the case of enantiomeric compounds, there are weak specific interactions between Br and N protonated atoms, with N-Br distance shown in table S2, SI.

The spacing between the inorganic PbBr₃⁻ arrays in the unit cell has been measured for all the compounds and the distances are

reported in Table 1. They are increasing in accordance with the size of the used cation when going from **BuA** to **MBA** then to **EBA**.

	Distance between two arrays (Å) c/2 and b (or a)
(<i>R</i>)- BuA -PbBr ₃	8.905 and 7.954
(<i>S</i>)- BuA -PbBr ₃	8.921 and 7.968
<i>rac</i> - BuA -PbBr ₃	8.812 and 7.971
(<i>R</i>)- MBA -PbBr ₃	10.131 and 7.886
(<i>S</i>)- MBA -PbBr ₃	10.148 and 7.902
(<i>R</i>)- EBA -PbBr ₃	10.250 and 8.174
(<i>S</i>)- EBA -PbBr ₃	10.2620 and 8.177
<i>rac</i> - EBA -PbBr ₃	10.651 and 8.019

Table 1: Distances between the 1D lead bromide units for (*S*)- or (*R*)-**BuA**-PbBr₃, **MBA**-PbBr₃, **EBA**-PbBr₃, (*rac*)-**BuA**-PbBr₃ and (*rac*)-**EBA**-PbBr₃

The analysis of the distortions around the Pb²⁺ cations in the 1D compounds has been performed using the Shape program.^{38,39} The octahedral distortion determined using a PbBr₆ O_h model is high (with CShM (Continuous Shape Measures) values varying between 14.013 and 19.070 (see ESI)). These values are high when compared to other values provided for 2D lead halides perovskites,⁴⁰ but we can assume a more distorted environment for 1D inorganic compounds. A careful analysis of the Pb-Br distances (see table S2 ESI) shows that a shorter Pb-Br distance (varying between 3.2382(10) and 3.4549(13)Å, for the 8 considered compounds) is observed for the Pb²⁺ cations in compounds presenting the highest CShM values towards octahedral PbBr₆ environment ((*S*)-**BuA**-PbBr₃, (*R*)-**BuA**-PbBr₃, *rac*-**BuA**-PbBr₃). The series of compounds with the **BuA** cation present the smaller deviation and the series with **EBA** and **MBA** cations present higher deviations. Interestingly, considering a PbBr₅ environment around the metal cations, (C_{4v} point group) as the undistorted structure (square pyramidal or truncated octahedron geometry), the calculations indicate that the series associated with the cations **EBA** and **MBA** present the smallest deviations. We can conclude that (*S*)-**BuA**-PbBr₃, (*R*)-**BuA**-PbBr₃ and *rac*-**BuA**-PbBr₃ present a strongly distorted octahedral geometry whereas (*S*)-**MBA**-PbBr₃, (*R*)-**MBA**-PbBr₃, (*S*)-**EBA**-PbBr₃, (*R*)-**EBA**-PbBr₃ and *rac*-**EBA**-PbBr₃ present a geometry closer to square pyramidal.

The resulting polycrystalline powders were analyzed using PXRD for all the compounds. The PXRD diagrams of all the polycrystalline samples evidence a good match between the simulated and experimental patterns, as shown in figures S1, confirming that the microcrystalline powders do not exhibit traces of impurities or other crystalline phases.

Optical characterization

The reflectance-UV spectra for the pairs of (*S*) and (*R*) **BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ enantiomers (figure S2) display an optical band gap lying between 2.8 and 3.4 eV.

The photoluminescence (PL) spectra measured under 325 nm (3.815 eV) continuous-wave excitation on large (mm-sized) crystals of the *R*- and *S*- enantiomer based compounds are shown in figure 4 a). All samples show a very broad emission spectrum with a maximum at around 1.9-2.0 eV, corresponding to a significant Stokes shift of above 1 eV, consistent with previous reports for STE emission^{19,23} and earlier observations for the **MBA**-PbBr₃ compound.³⁴ The excitonic character of the broad emission band is confirmed by the linear dependence of the integrated intensity on the excitation power for all of the investigated compounds, in agreement with what has previously been reported for STE emission.^{23,19} In contrast, a saturation at high excitation intensity would be expected for defect-related PL features. Figure 4 b) shows the excitation power dependence of the integrated emission intensity for the (*R*) enantiomers of the different compounds. The full set of data is presented in the Supporting Information, figure S3, ESI. This allows to rule out defect-related character of the broad PL band⁴¹, for which a sub-linear dependence should be observed. It should be noticed that no changes of spectral shape can be observed when increasing the excitation power (see Supporting Information, Figure S4, ESI), ruling out donor-acceptor pair emission or defect saturation effects.

The maximum of the emission is slightly spectrally shifted for the different compounds and a shoulder at around 2.5 eV is observed only for the **EBA**-PbBr₃ crystal. It is possibly related to a local impurity although we cannot exclude the presence of a different emissive self-localized state at higher energy.⁴² The Stokes-shifted emission was dominant and no significant signal related to the recombination of FE (Free Excitons) could be well resolved. The presence of an energy barrier separating the FE and STE states can explain the quenching of the FE emission and the presence of only STE emission at elevated temperatures.^{21,23} The observation of a broadband STE emission at room temperature could be correlated with the observed octahedral distortion. In 2D hybrid perovskite, the intensity of the STE emission relative to the FE emission is found to increase linearly with the out of plane distortion of the Pb-(μ-X)-Pb (X = Cl-/Br-) angle in the inorganic sheet.⁴³ It has been established that the broadening and the intensity of the photoluminescence emission spectrum of low dimensional HOIPs (mainly 2D) are strongly related with the distortion around the metal cations⁴⁰ and also in the inorganic lattice.⁴³

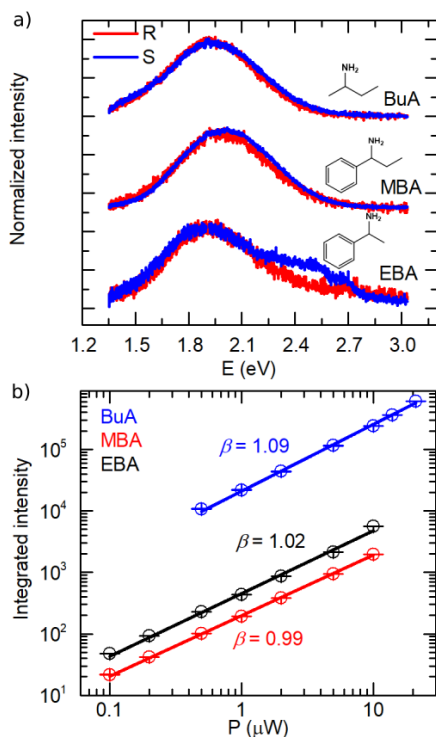


Figure 4: a) PL emission spectra of (S)- or (R)-**BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ compounds measured using 325 nm (3.815 eV) excitation. b) Integrated intensity of the broad emission peak as a function of the 3.06 eV excitation power for the (R) enantiomers of the three compounds. The experimental data is marked in blue (**BuA**-PbBr₃), red (**MBA**-PbBr₃) and black (**EBA**-PbBr₃) symbols. The lines show the fit of the power law $I = aP^\beta$.

To further characterize the broadband emission, the Commission Internationale de l'Éclairage (CIE) chromaticity coordinates of enantiomerically pure compounds were determined. The CIE chromaticity coordinates (figure 5) were determined as (0.53, 0.45) for **BuA**-PbBr₃, (0.50, 0.46) for **MBA**-PbBr₃ and (0.49, 0.43) for **EBA**-PbBr₃, corresponding to a larger contribution in the orange-red region of the spectrum. The **Correlated Color Temperatures (CCT)** were determined as 2238 K, 2620 K and 2404 K for the **BuA**-, **MBA**- and **EBA** compounds, respectively. Broadband emission at such low energy spectral range has recently been reported for other 1D perovskite compounds^{44,45} as well as for 0D metal halides⁴⁶ and 0D mixed halide perovskites [Zhou] and is considered highly interesting for applications in light-emitting devices [Feng] [Zhou].^{44,47}

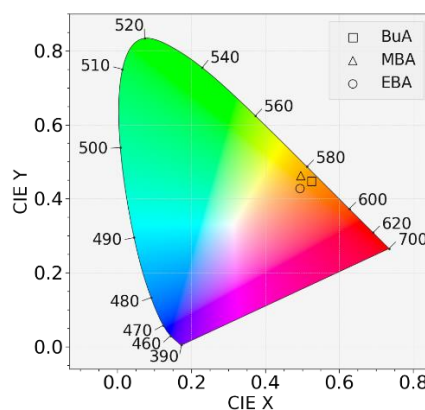


Figure 5: The CIE coordinates for (S)- or (R)-**BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ compounds. Because of the spectral similarity of the emission for the (R) and (S) enantiomers, only the coordinates calculated for the (R)-enantiomer compounds are shown for better clarity.

Comparing the PL emission spectra obtained using 3.18 eV and 3.06 eV excitation (Fig. S4) we observed a blueshift of the broadband emission for lower excitation energy, which is consistent with previous observations for STE emission in low-dimensional perovskites and could indicate the existence of multiple STE states with different energies [23] [Zhao].

We have also investigated the polarization properties of the diffuse reflectance and PL emission. The diffuse reflectance circular dichroism (DRCD) spectra for the pairs of (S) and (R)-**BuA**-PbBr₃, **MBA**-PbBr₃ and **EBA**-PbBr₃ enantiomers are presented in figure S5. They attest the presence of enantiomers in the crystalline phase of each compound.

We then analyzed the degree of linear polarization of the broadband emission for the (R)-**BuA**-PbBr₃, (S)-**BuA**-PbBr₃ and (rac)-**BuA**-PbBr₃ compounds. Due to the 1D structure of the compounds, comparable to the structure of quantum wires, a significant anisotropy of the PL emission can be expected.

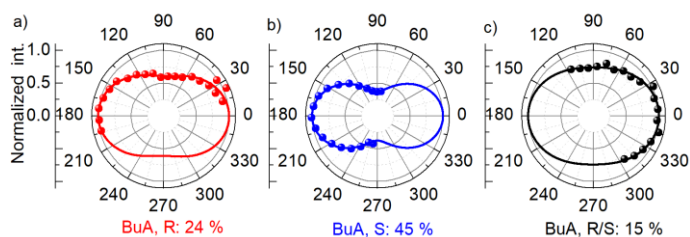


Figure 6: Polar plots of the normalized intensity of the photoluminescence emission as a function of detection angle measured for 3.815 eV excitation for a) (R)-**BuA**-PbBr₃, b) (S)-**BuA**-PbBr₃ and c) (rac)-**BuA**-PbBr₃.

Figure 6 shows the normalized integrated emission intensity as a function of the orientation of the linear polarization analyzer. The data can be fitted by the Malus' law: $I(\theta) = I_0 + A \cos^2(\theta - \theta_0)$ and the DLP (**Degree of Linear Polarization**) is defined as $A/(A+2I_0)$. The DLP is 24% for (R)-**BuA**-PbBr₃, 45%

for the (*S*)-**BuA**-PbBr₃ and 15 % for (*rac*)-**BuA**-PbBr₃, which is significant for a bulk crystal at room temperature.¹⁸

No detectable circular polarization of the PL could be measured, once we exclude any artifacts that could arise due to the presence of linear polarization.⁴⁸ Enhanced carrier-phonon interaction and thermal lattice expansion at elevated temperatures, as well as increased probability of a spin-flip may explain the absence of circular polarization at room temperature.⁴⁹

High degrees of linear polarization can be achieved for PL emission from individual nanowires and nanocrystals.^{50,51,52,53} However, to achieve polarized emission at the macroscopic scale, the synthesis of composite structures is generally required, such as nanocrystals embedded in stretched polymer films (polarization ratio of around 30 %) ⁵⁴ or dispersed in nematic liquid crystals (PL anisotropies up to 30 %).⁵⁵ Here, we observe a significant linear polarization from macroscopic crystals which can be explained by the intrinsic anisotropy of the transition dipole moments and quantum confinement effects as well as additional anisotropic distortions of the octahedral of the inorganic lattice responsible for the self-trapping in the 1D structure.⁴³ The presence of significant DLP in macroscopic single crystals proves good crystalline quality since the presence of differently oriented crystalline domains would lead to the loss of correlation between the macroscopic long axis of a crystal and the linear polarization direction.¹⁸ In this context, our observation of strong linear polarization in macroscopic crystals is interesting from the point of view of potential applications in polarized light emission^{56,57} and detection.⁵⁸

Conclusions

In this work, we investigated the formation of a series of chiral 1D perovskites built from PbBr₃ and ammonium based enantiomerically pure cations (**BuA**, **MBA** and **EBA**) inserted between the 1D inorganic arrays. For **BuA** and **EBA**, the analogous racemic compounds were obtained, and they present the same structural arrangement as for the pure enantiomers. Broadband below-gap photoluminescence was observed for all of the synthesized perovskite compounds and attributed to the recombination of self-trapped excitons. The yellow to orange spectral range of the STE emission makes the investigated compounds highly interesting for light-emitting devices such as optically pumped LEDs, where they could be used in combination with blue-emitting phosphors to obtain a high color rendering index and overcome the issue of deficiency in red emission encountered for many yellow phosphors.⁴⁷

For all enantiomers of the **BuA**-based compound, polarization-resolved PL emission measurements evidence a significant linear polarization at RT as the consequence of the structural anisotropy of the 1D crystal structure. Broadband polarized light sources are highly in demand for applications for display, imaging and optical communications devices⁵⁹ and the compounds described in this work expand the new family of

polarized broadband light-emitting perovskite materials. Our results provide an example of how the structural distortions and the resulting photophysical properties can be driven by a targeted choice of the organic cations in low-dimensional HOIPs.

Experimental

Synthesis

All chemicals and reagents (starting amine) were of analytical grade and used as received from commercial sources (Sigma-Aldrich). The solvent used in these syntheses was distilled water.

Crystals of (*S*)-**BuA**-PbBr₃, (*R*)-**BuA**-PbBr₃ and (*rac*)-**BuA**-PbBr₃.

To a solution of 0.752g (2.05mmol) of PbBr₂ in 4ml of a 48% HBr solution, 0.3g (4.1mmol) of (*R*)-, (*S*)- or (*rac*)-**BuA** were added. The mixture was stirred at room temperature for 1h. Ethanol solution was diffused by evaporation into the mixture. Many colorless crystals produced after 24h were collected and washed by Ethanol and diethyl ether.

Crystals of (*S*)-**MBA**-PbBr and (*R*)-**MBA**-PbBr₃

To a solution of 0.757g (2.06mmol) of PbBr₂ in 6ml of a 48% HBr solution, 0.5g (2.35mmol) of (*R*)- or (*S*)-**MBA** were added. Then the mixture was heated at 80°C for 1h. The clear solution was slowly cooled to room temperature which leads to colorless needles. The crystals were collected and washed by ethanol and diethyl ether.

Crystals of (*S*)-**EBA**-PbBr₃, (*R*)-**EBA**-PbBr₃ and (*rac*)-**EBA**-PbBr₃.

Colorless single crystals were all grown following the procedure described above using the same molar ratios, starting from (*R*)-, (*S*)- or (*rac*)-**EBA**.

X-Ray data collection

Data were collected at 173(2) K on a Bruker APEX8 CCD Diffractometer equipped with an Oxford Cryosystem liquid N₂ device, using graphite-monochromated Mo-K α (λ = 0.71073 Å) radiation. For both structures, diffraction data were corrected for absorption. Structures were solved using SHELXS-97 and refined by full matrix least-squares on F² using SHELXL-97. The hydrogen atoms were introduced at calculated positions and not refined (riding model). They can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/datarequest/cif. CCDC: (*S*)-**BuA**-PbBr₃ (2167654), (*R*)-**BuA**-PbBr₃ (2167649), *rac*-**BuA**-PbBr₃ (2167653), (*S*)-**MBA**-PbBr₃ (2167650), (*R*)-**MBA**-PbBr₃ (2167652), (*S*)-**EBA**-PbBr₃ (2167655), (*R*)-**EBA**-PbBr₃ (2167656), *rac*-**EBA**-PbBr₃ (2167651).

Structural deviations calculations

The structural deviations calculations leading to the obtention of the CShM (Continuous Shape Measures) values have been obtained using the Shape program.³⁸ Based on a general polyhedral approach, computational details are provided in the literature.³⁹

Powder diffraction studies (PXRD)

Diagrams were collected on a Bruker D8 diffractometer using monochromatic Cu-K α radiation with a scanning range between 4 and 40° using a scan step size of 8° mn⁻¹.

Absorbance

The absorption spectra were measured in the solid-state using a PE Lambda 650 S spectrophotometer at room temperature.

Circular Dichroism

Diffused reflectance circular dichroism (DRCD) spectra of crystalline samples were performed on a JASCO J-1500 spectrophotometer equipped with an integrating sphere attachment (DRCD-575 accessory) in the 200-400 nm domain at a scan rate of 100 nm/min and accumulated twelve times.

The solid samples were diluted with KBr to a concentration of 25% (w/w).

Photoluminescent measurements

All the presented optical measurements were performed under ambient conditions. The photoluminescence spectra measured using 3.815 eV excitation, including the polarization-resolved data, were recorded in a macro-PL setup at 45-degree incidence angle using the a HeCd continuous wave laser for excitation. The signal was collected, dispersed by a spectrometer and detected using a CCD camera. For polarization-resolved measurements, an analyser composed of a waveplate and a linear polarizer was placed in front of the spectrometer in the detection path and the waveplate was rotated to detect different polarization

states of the emitted light. We used a half-waveplate for the linear polarization measurements and a quarter waveplate for the circular polarization detection. Based on the analysis of the emission of an unpolarized light source, we determined the precision of our polarization detection to be around 3% accounting for instrumental error and the excitation beam intensity fluctuations.

Measurements using 3.06 eV (405 nm) pulsed picosecond excitation, including power-dependent series, were performed in a micro-PL setup in backscattering configuration. The measurements were performed at 5 MHz excitation repetition rate and the size of the spot in the micro-PL configuration is on the order of 1 μm^2 .

Since we do not perfectly control the in-plane orientation of the crystalline axis with respect to the detection plane it is possible that the measured DLP is only a lower bound.²⁰

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Financial support from the University of Strasbourg and the CNRS is acknowledged. Dr. Catalin Maxim (University of Bucharest, Romania) is acknowledged for recording reflectance-CD spectra. The work of J.M. Urban, D. Garrot, G. Trippé-Allard and E. Deleporte is financially supported by Agence Nationale de la Recherche: EMIPERO Project ANR-18-CE24-0016.

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